

Equilibrium Expressions

Equilibrium Expressions – manipulating and using equilibrium constants

Equilibrium constants have been measured for many reactions. Since in the equilibrium expression given by the "[Law of Mass Action](#)" the equilibrium constant K is a ratio of product concentrations over reactant concentrations, large K implies products will be largely present while small K implies reactants will be largely present when equilibrium has been reached. Even more usefully, if initial concentrations (and K) are known, then equilibrium concentrations can be calculated with elementary algebra.

From Perdue here are the [basics](#) and the [rules](#) for [writing equilibrium expressions](#). And [here](#) is a review on [how to calculate equilibrium constants](#) from experimental data.

Hyperlinks on these pages are borrowed from various academic institutions and are included to help you consolidate your chemical knowledge.